



# (CuInSe<sub>2</sub>)<sub>1-x</sub> (TaSe)<sub>x</sub> Solid Solutions (0 < X ≤ 0.5): X-Ray Diffraction, Scanning Electron Microscopy, Differential Thermal Analysis and Magnetic Susceptibility

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## Abstract

Polycrystalline samples of the (CuInSe<sub>2</sub>)<sub>1-x</sub> (TaSe)<sub>x</sub> system were prepared by the melt and anneal technique in the composition range 0 < x ≤ 0.5. Products were characterized by X-Ray Diffraction (XRD), Scanning Electron Microscopy (SEM), Differential Thermal Analysis (DTA) and SQUID techniques. From XRD and SEM, it was found that the solubility of TaSe in CuInSe<sub>2</sub> is around 10%; however, samples up to x = 0.5 are composed by a mean tetragonal CuInSe<sub>2</sub>-like phase with traces of TaIn<sub>0.67</sub>Se<sub>2</sub>. A preliminary T-x phase diagram is presented. DC magnetic susceptibility measurements as a function of temperature using the ZFC-FC protocol were performed in samples x = 0.05 and 0.5. The magnetic behavior evolves from paramagnetic to weak ferromagnetic. In both samples, the presence of hysteresis between FC and ZFC curves suggest the presence of magnetic clusters; fitting FC curves with a Langevin function it was found that clusters are composed for ~10<sup>3</sup> Ta-atoms.

## Introduction

CuInSe<sub>2</sub> is an A<sup>I</sup>-B<sup>III</sup>-C<sup>VI</sup><sub>2</sub> semiconductor material which crystallize in the tetragonal chalcopyrite structure, space group  $I\bar{4}2d$  (No122), widely studied because of its applications in the photovoltaic industry [1-5]. Moreover, the discovery in the last years of room tem-

perature ferromagnetism (RT-FM) in the analogous A<sup>II</sup>-B<sup>IV</sup>-C<sup>V</sup><sub>2</sub> compounds doped with transition metals (TM) [6-22] has prompted a new and renewed interest in his research. Particularly, the magnetism of CuInSe<sub>2</sub> doped and/or alloyed with Mn, Fe and Ta has been reported before [18,19,23,24] founding paramagnetic, superparamagnetic and spinglass behaviors (Table 1).

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**Table 1:** CuInSe<sub>2</sub>: TM alloys previously reported in the literature.

Alloys	Magnetic element composition	Synthesis method	Magnetic behavior	T <sub>c</sub> [K]	Reference
Cu <sub>1-x</sub> Mn <sub>x</sub> InSe <sub>2</sub>	0 < x < 0.2	SSR	PM	---	[18]
CuIn <sub>1-x</sub> Mn <sub>x</sub> Se <sub>2</sub>	0.0125 < x < 0.2	SSR	PM	---	[19]
Cu <sub>1-y</sub> In <sub>1-y</sub> Mn <sub>2y</sub> Se <sub>2</sub>	0.0125 < y < 0.6				
(CuIn) <sub>1-x</sub> Fe <sub>x</sub> Se <sub>2-x</sub>	x = 0.5	SSR	SPM	---	[23]
(CuIn) <sub>1-x</sub> Ta <sub>x</sub> Se <sub>2-x</sub>	x = 0.25	SSR	SG	50	[24]

RT-FM arise from the interaction of holes (created by substitution of B<sup>IV</sup> or B<sup>III</sup> cations by TM) with the local moment of the TM *d* electrons [25,26]; however, it is also argued that it is due by the presence of magnetic secondary phases such as MnP and MnAs [26-30]. Kochura, et al. in samples prepared by solid state reaction under the condition of fast cooling, found that exist three types of magnetic species in A<sup>II</sup>B<sup>IV</sup>C<sup>V</sup><sub>2</sub>: Mn alloys: a) Substitution of Mn ions making Mn complexes (especially dimers); b) MnAs micro-precipitates and c) MnAs nanosize precipitates (clusters with a mean diameter of 3 nm) [27].

Although the origin of RT-FM is still matter of investigation, it is well established that is composition dependent: at low values of x, the alloy shows a typical paramagnetic behavior (or also superparamagnetic) [9,11] whereas for higher values of x, a critical x<sub>c</sub> value is attained for which, the paramagnetic→ferromagnetic transition occurs. It is also worth to note that T<sub>c</sub> values are approximately the same for all A<sup>II</sup>B<sup>IV</sup>C<sup>V</sup><sub>2</sub>: Mn alloys suggesting that this transition is related to Mn-based secondary phases more than substitution of cations for Mn<sup>2+</sup> in the ternary matrix.

It is interesting to compare theoretical calculations with experimental results. Katamani, et al. [31,32], using KKR-CPA-LDA method, predict that ferromagnetic states are stable in (Cd<sub>1-x</sub>V<sub>x</sub>)GeP<sub>2</sub> and (Cd<sub>1-x</sub>Cr<sub>x</sub>)GeP<sub>2</sub> alloys, whereas (Cd<sub>1-x</sub>Mn<sub>x</sub>)GeP<sub>2</sub>, (Cd<sub>1-x</sub>Fe<sub>x</sub>)GeP<sub>2</sub> and (Cd<sub>1-x</sub>Co<sub>x</sub>)GeP<sub>2</sub> must show spinglass-like ground states (calculations were made using x = 0.1). They also found that Ti and Ni substitutions in CdGeP<sub>2</sub> could not have a net magnetic moment (the same are applicable to ZnGeP<sub>2</sub> and CdSiAs<sub>2</sub>). On the other hand, the ferromagnetic state was found to be stable in AgGaS<sub>2</sub> (CuAlS<sub>2</sub>) doped with Ti, V, Cr and Mn, whereas doping with Fe, Co and Ni must stabilizes a spinglass-like state. The work of Zhao, et al. [25], using first principle calculations, coincides with Katamani and predicts that Mn doping at the III site in A<sup>I</sup>B<sup>III</sup>C<sup>VI</sup><sub>2</sub> compounds provides holes that must stabilize ferromagnetic coupling between Mn ions. The predictions of both works (Katamani and Zhao) are contradictory that which has been observed experimentally! The reason for the discrepancies between theoretical expectations and experimental results is not clear until now. It has been suggested that stabilization of ferromagnetism is also due to the interaction of holes with Mn

magnetic moments induced by:

- Intrinsic defects [33,34];
- Non-uniform spatial distribution of carriers and/or and magnetic ions [35]; or
- Nanoscale phase separation, driven either by randomness in the carrier and spin subsystems or by limited solubility of transition metals in the host semiconductor, which leads to spinodal decomposition into regions with a small and a large concentration of the magnetic constituents [35].

It is evident that additional experimental work, in greater variety of systems and with larger concentration of the transition metal dopant, is necessary.

In the present work we report the preparation and characterization of the (CuInSe<sub>2</sub>)<sub>1-x</sub>(TaSe)<sub>x</sub> solid solutions system in the composition range 0 < x ≤ 0.5, as a part of our systematic research in A<sup>I</sup>B<sup>III</sup>C<sup>VI</sup><sub>2</sub> alloys with TM. Some partial results had been previously reported: XRD diffraction and DTA were presented as short note in a conference [36], for sample x = 0.5 (CuTaInSe<sub>3</sub>) the preparation and crystal structure [37,38] and for sample x = 0.25 the magnetic susceptibility [39]. Here, in this work, we are added a more exhaustive analysis of XRD diffraction, identification of the phases by SEM measurements, a preliminary T-x phase diagram, and DC magnetic susceptibility measurements for samples x = 0.05 and x = 0.5

## Experimental Procedure

### Preparation

Samples were prepared in steps of x = 0.1 in the composition range 0 < x ≤ 0.5 and in addition samples x = 0.05 and 1/4 were also prepared. Polycrystalline ingots, of about 1 g were prepared by the usual melt and anneal technique. Starting materials (Cu, In, Ta, and Se) with a nominal purity of 99.99 wt. % in the stoichiometric ratio were mixed together in an evacuated (10<sup>-4</sup> Torr) and sealed quartz tube with the inner walls previously carbonized in order to prevent chemical reaction of the elements with quartz. The quartz ampoule is heated until 493 K (melting point of Se) keeping this temperature for 48 h and shaking all the time by means of a electromechanical motor. This procedure guarantees the forma-

tion of binary species at low temperature avoiding the existence of Se free gas at high temperature, which could produce explosions or a Se deficiency in the ingot. Then the temperature is slowly increased until 1423 K, with the mechanical shaker connected for a better mix of the components. After 24 h, the cooling cycle begins until the anneal temperature (800 K) with the mechanical shaker disconnected. The ampoule is keeping at the annealing temperature during 1 month in order to assure the thermal equilibrium. Then the furnace is switching off.

### Scanning Electron Microscopy (SEM)

The stoichiometric relation of the samples was investigated by scanning electron microscopy (SEM) technique, using Hitachi S2500 equipment. The micro chemical composition was found by an energy-dispersive x-ray spectrometer (EDS) coupled with a computer-based multichannel analyzer (MCA), (Delta III analysis and Quantex software, Kevex). For the EDS analysis,  $K_{\alpha}$  lines were used. Accelerating voltage was 15 kV. The samples were tilted 35 degrees. A standarless EDS analysis was made with a relative error of  $\pm 5$ -10% and detection limits of the order of 0.3 wt. %, where the k-ratios are based on theoretical standards.

### X-Ray Diffraction (XRD)

X-ray powder diffraction data was collected by means of a diffractometer (Bruker D5005) equipped with a graphite monochromator ( $\text{CuK}_{\alpha}$ ,  $\lambda = 1.54059 \text{ \AA}$ ) at 40 kV and 20 mA. Silicon powder was used as an external standard. The samples were scanned from  $10$ - $100^{\circ} 2\theta$ , with a step size of  $0.02^{\circ}$  and counting time of 20 s. The Bruker analytical software was used to establish the positions of the peaks from the  $\text{CuK}\alpha_1$  component and to strip mathematically the  $\text{CuK}\alpha_2$  components from each reflection. The peak positions were extracted by means of single-peak profile fitting carried out through the Bruker DIFFRAC<sup>plus</sup> software. Each reflection was modeled by means of a pseudo-Voigt function.

### Differential Thermal Analysis (DTA)

The differential Thermal Analysis (DTA) was carried out in a fully automatic Perkin-Elmer apparatus with Pt/Pt-Rh thermocouples. Au was used as internal standard. The heating and cooling rates were controlled to  $20 \text{ Kh}^{-1}$ . Transition temperatures were manually obtained from the  $\Delta T$  vs. T graph with the criteria that the transition occurs at the intersection of the base line with the slope of the thermal transition peak, as usually. The maximum error committed in the determination of transition temperatures by this method was estimated at  $\pm 10 \text{ K}$ .

### Magnetic susceptibility

DC magnetic susceptibility as a function of tempera-

ture measurements,  $\gamma(T)$ , were performed on a Quantum Design SQUID magnetometer, equipped with a superconducting magnet able to produce fields up to 5 T. The samples in the form of powder were compacted with a piece of cotton inside the sample holder in order to prevent any movement. Zero-field-cooling and field cooling (ZFC-FC) measurements were carried out in the temperature range of 2-300 K.

### Analysis and Discussion

As it is routine in our laboratory, stoichiometric relations of all the samples were checked out in a thin slide ( $\sim 1 \text{ mm}$  thick) cut off from the center of the ingot. The results are show in Table 2. The experimental error is estimated in the first decimal place ( $\pm 0.1$ ). The sum of the percentages does not always add exactly 100% by approximations of decimals.

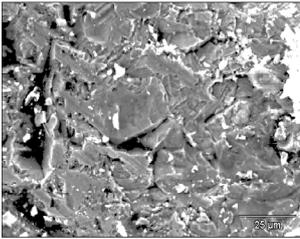
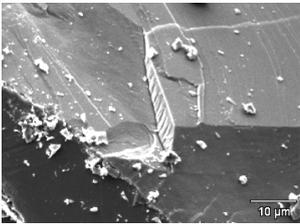
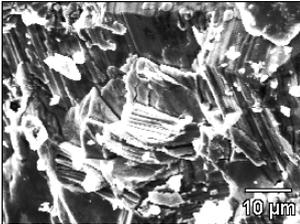
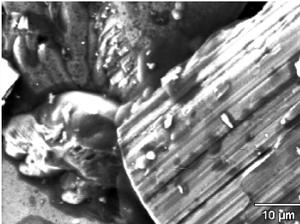
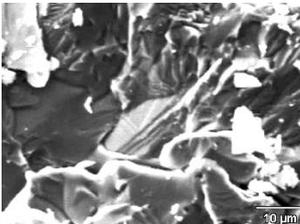
General observations are:

- Deviations in the stoichiometry for Cu and Se, in average, lies inside of the estimated experimental error ( $\sim 10\%$ );
- In the case of In and Ta, deviations (always in average) are higher that the experimental error (20 and 30%, respectively);
- In average, the quantity of Ta in the samples does not excess 10%, suggesting that this is the limit of solid solubility.

In particular, in sample  $x = 0.1$ , Cu copper stoichiometry is close to nominal, whereas In and Ta are somewhat low. For sample  $x = 0.2$ , three experiences give a similar behavior to sample  $x = 0.1$  but two experiments (2 and 3) give lower stoichiometry values for Cu and In and higher for Ta with respect to nominal. Behavior of sample  $x = 0.3$  is similar to sample  $x = 0.2$ . For samples  $x = 0.1, 0.2$  and  $0.3$ , although the nominal stoichiometry values of Cu and In are the same, the experimental values of Cu are always higher than In; this behavior is inverted for sample  $x = 0.4$ . Sample  $x = 0.5$  have a similar behavior of samples  $x = 0.1, 0.2$  and  $0.3$  with respect to the relation between Cu and In, but the Ta values are too lower with respect to nominal, with exception of exp. 1 where the value of Ta is higher than nominal but of Cu and In are lower.

With respect to the distribution of Cu, In Ta and Se atoms in the samples, it can be observed in Figure 1 that it is quite homogenous (Cu: red, Se: yellow, Indium: blue, Tantalum: green); this indicative that our observations are more due to local fluctuations of stoichiometry that large regions of different phases. Effectively, in SEM technique the usual resolution is in the order of one digit in nanometer units (although in some sophisticat-

**Table 2:** Microphotographs and measured stoichiometry for  $(\text{CuInSe}_2)_{1-x}(\text{TaSe})_x$  alloys with  $x = 0.1, 0.2, 0.3, 0.4$  and  $0.5$ . In red the nominal stoichiometry.

Microphotographs	Nominal Composition	Cu at%	In at%	Se at%	Ta at%
	<b>x = 0.1</b>	<b>23.7</b>	<b>23.7</b>	<b>50.0</b>	<b>2.6</b>
	Exp. 1	22.7	17.7	58.8	0.8
	Exp. 2	21.7	18.1	59.4	0.8
	Exp. 3	22.1	18.4	58.3	1.2
	Exp. 4	21.1	16.7	60.4	1.7
	Exp. 5	21.9	17.8	59.2	1.2
	<b>x = 0.2</b>	<b>22.2</b>	<b>22.2</b>	<b>50.0</b>	<b>5.6</b>
	Exp. 1	20.2	18.1	59.1	2.5
	Exp. 2	16.4	13.5	61.1	9.0
	Exp. 3	16.4	12.5	62.1	9.0
	Exp. 4	20.1	16.1	60.7	3.1
	Exp. 5	20.2	17.1	59.9	2.8
	<b>x = 0.3</b>	<b>20.6</b>	<b>20.6</b>	<b>50.0</b>	<b>8.8</b>
	Exp. 1	16.8	12.5	60.2	10.5
	Exp. 2	14.8	11.9	61.0	12.3
	Exp. 3	22.3	15.6	53.2	8.7
	Exp. 4	28.7	17.3	55.6	6.4
	Exp. 5	20.7	14.4	57.5	9.5
	<b>x = 0.4</b>	<b>18.8</b>	<b>18.8</b>	<b>50.0</b>	<b>12.5</b>
	Exp. 1	12.5	19.5	54.6	12.5
	Exp. 2	16.8	18.6	57.3	7.3
	Exp. 3	17.3	20.1	57.6	5.0
	Exp. 4	18.8	21.1	58.0	2.1
	Exp. 5	16.7	19.8	56.9	6.7
	<b>x = 0.5</b>	<b>16.7</b>	<b>16.7</b>	<b>50.0</b>	<b>16.7</b>
	Exp. 1	11.8	3.6	65.6	19.1
	Exp. 2	18.8	14.3	59.0	7.9
	Exp. 3	21.3	13.2	61.3	4.3
	Exp. 4	22.3	16.0	59.7	2.1
	Exp. 5	20.8	14.5	60.1	4.6

ed equipment it can be obtained resolutions of 0.4 nm [40]), that is  $\sim 10^{-4}$  times the size of the white trait scale in Figure 1.

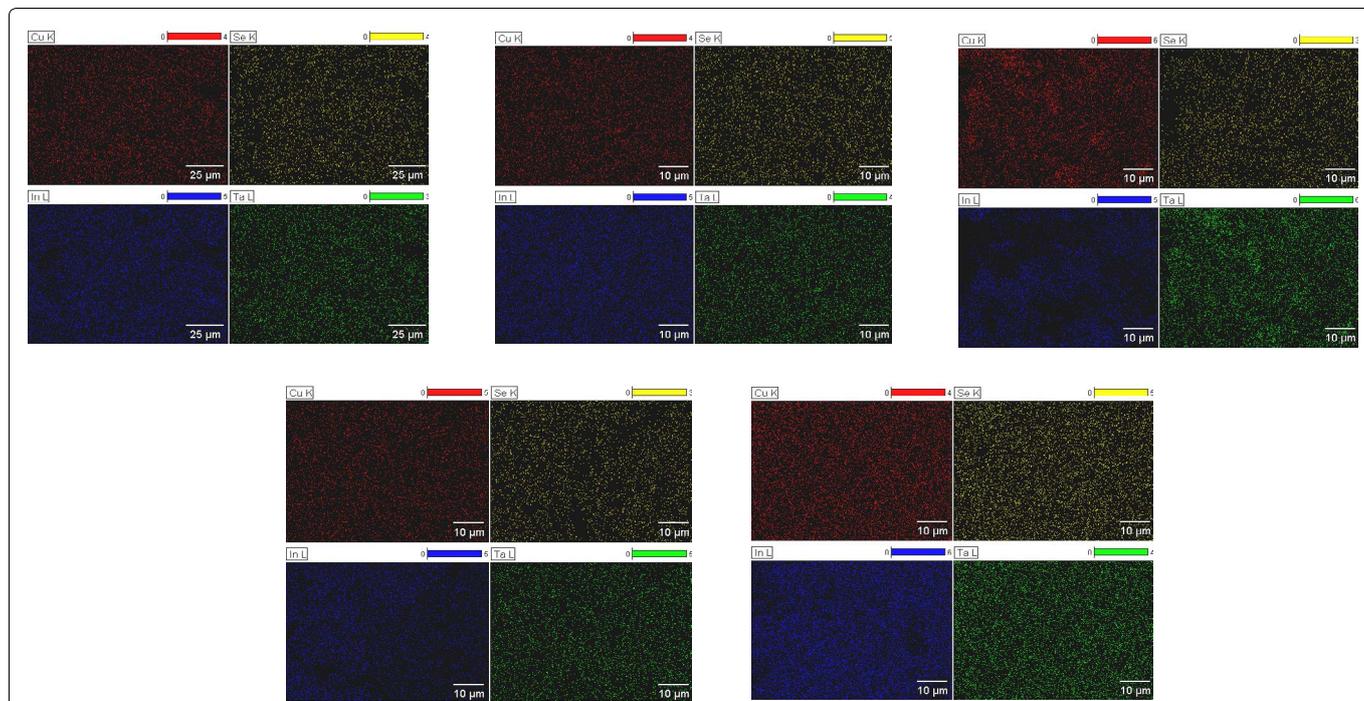
However, under a more exhaustive examination, for sample  $x = 1/3$ , we are observed little structures with hexagonal symmetry (secondary phase) which we are identified as  $\text{TaIn}_{0.67}\text{Se}_2$  (Figure 2).  $\text{TaIn}_{0.67}\text{Se}_2$  is a previously reported compound [41], crystallizes in a hexagonal structure, space group  $P\bar{6}m2$  (N 187) with cell parameters  $a = b = 3.45 \text{ \AA}$ ,  $c = 8.22 \text{ \AA}$ ,  $\alpha = \beta = 90^\circ$ ,  $\gamma = 120^\circ$ . The nominal stoichiometry for  $\text{TaIn}_{0.67}\text{Se}_2$  is Ta: 27.25%, In: 18.26% and Se: 54.50% which is very close to those observed in points 1, 3, 4 and 5 (left microphotography, upper table)

**Table 3:** Values of the covalent radii for Cu, In and Ta.

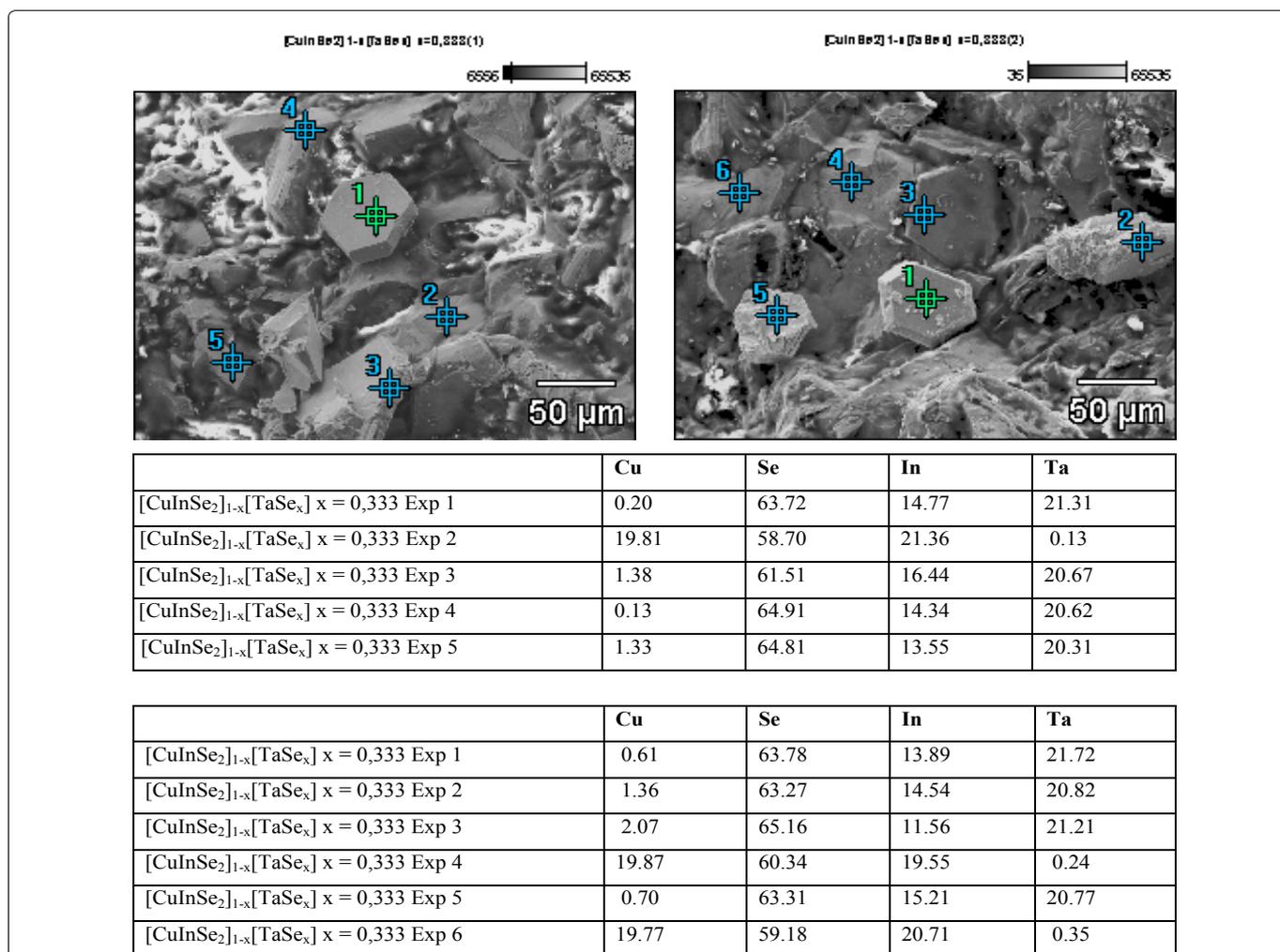
Cation	Covalent radius [Å]				
	Ref [42]	Ref [43]	Ref [44]	Ref [45]	Ref [46]
Cu	1.17	1.38	1.38	1.17	1.12
In	1.44	1.44	1.44	1.50	1.42
Ta	1.34	1.38	1.45	1.34	1.46

and 1, 2, 3 and 4 (right microphotography, bottom table). The hexagonal symmetry is clearly observed in microphotographs: point 1 of left microphotography and point 3 of right microphotography.

The lattice parameters of  $(\text{CuInSe}_2)_{1-x}(\text{TaSe})_x$  solid solutions were previously reported [37] (Figure 3), it can be observed that parameters  $a$  and  $c$  remains nearly con-

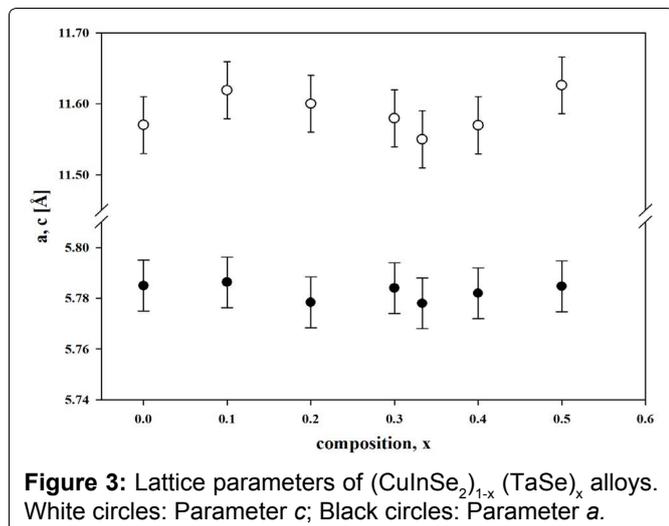


**Figure 1:** Distribution of Cu, Se, In and Ta in the samples  $x = 0.1, 0.2, 0.3, 0.4$  and  $0.5$ .



**Figure 2:** Microphotographs of sample  $x = 1/3$ . The tables show the atomic concentrations (%) for the experimental points (left photography → upper table and right photography → bottom table).

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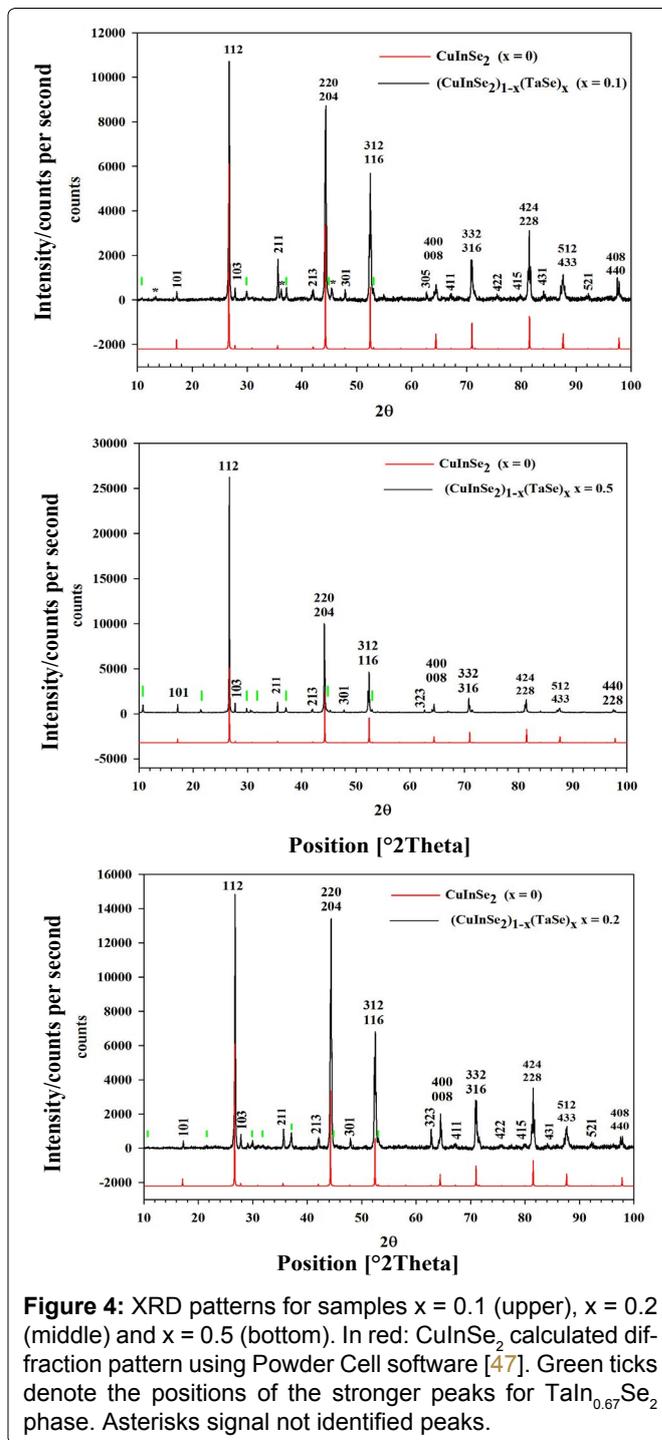


stant in the entire composition range studied. However, it is worth a short discussion about the covalent radii of Cu, Ta and In (Table 3) and the occupancy of the Ta atoms in the  $\text{CuInSe}_2$  matrix. Although there is a considerable dispersion in reference to the covalent radius values from different authors, it is well established that it increases from left to right and from upper to bottom in the periodic table: Cu has the lower value whereas Ta and In are similar. The fact that the lattice parameters do not change with composition in  $(\text{CuInSe}_2)_{1-x}(\text{TaSe})_x$  solid solutions up to  $x = 0.5$  suggests that Ta preferably substitutes In in the  $\text{CuInSe}_2$  matrix. That is, may be the reason that the observed general trend of deviation in stoichiometry with respect to the nominal values is higher in In than in Cu.

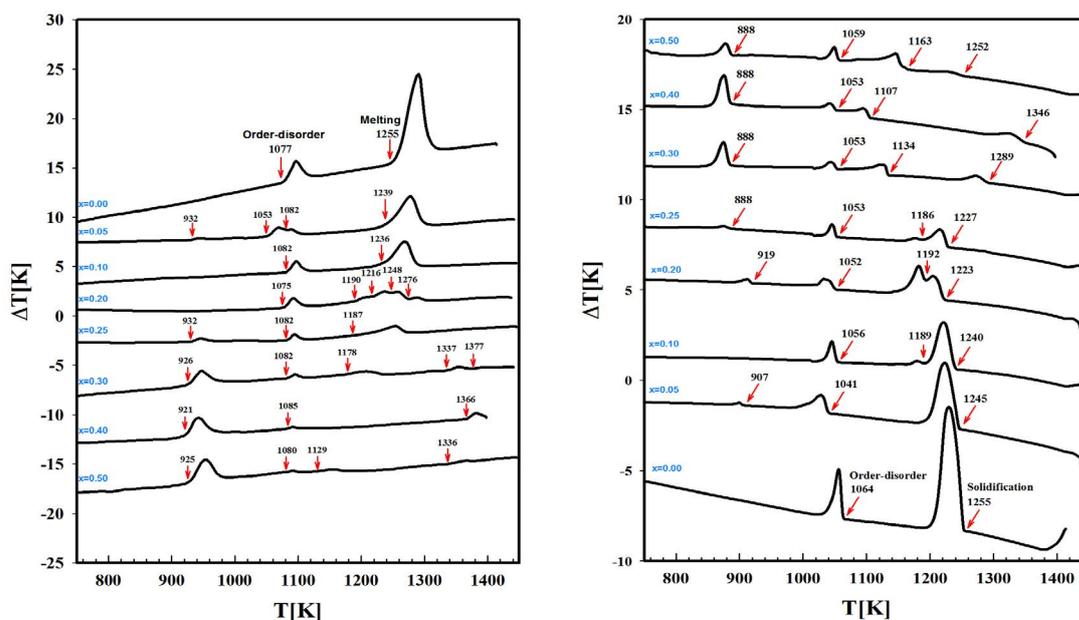
In Figure 4, we are displayed the diffraction patterns of compositions  $x = 0.1$ ,  $0.2$  and  $0.5$ , which have not showed in a previous report [37]. In this figure:

- The experimental diffraction patterns is displayed in black,
- A calculated diffraction pattern of  $\text{CuInSe}_2$ , using Powder Cell software [47] in red,
- Green ticks denote the stronger peaks of  $\text{TaIn}_{0.67}\text{Se}_2$  phase which was also observed in SEM measurements,
- Asterisks denotes not identified peaks.

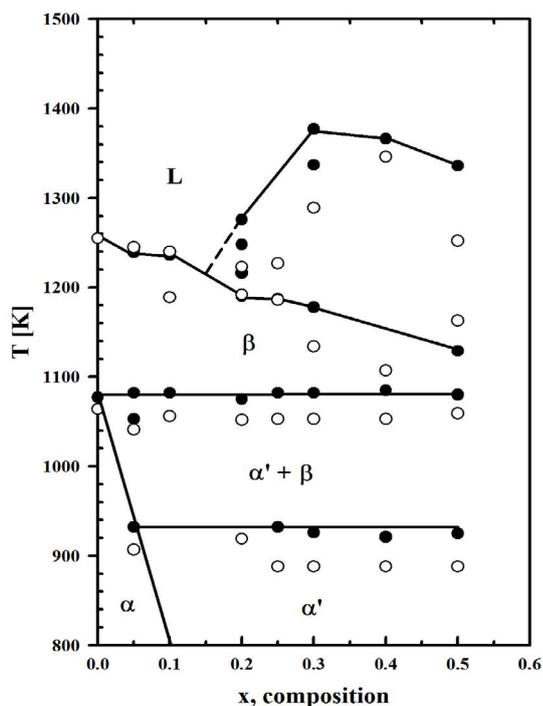
The coincidence of the  $\text{CuInSe}_2$  peaks with those of the alloys ( $x = 0.1$ ,  $0.2$  and  $0.5$ ) confirms that the lattice parameters remain constant in the entire composition range studied. The mean difference between diffraction patterns results in peak intensities. The intensities of the diffraction peaks are related with the occupation of the crystallographic sites by cations; however, Ta will replace no way of knowing, a priori, which cation, Cu or In. Moreover, refinements using Rietveld method suggest



that at least one crystallographic site is occupied at random for the three cations, as it is the case of  $\text{CuTaInSe}_3$ , the sample  $x = 0.5$  [24]. With respect to this sample ( $x = 0.5$ ) it can be observed in Figure 3 that the relation between the intensities of the two stronger peaks ( $I_{112}/I_{220}$ ) is 2.62, whereas for samples  $x = 0$ ,  $x = 0.1$  and  $x = 0.2$  the values are 1.53, 1.22 and 1.10, respectively. In previous works on chalcopyrite/metal transition alloys we are extensively discussed that the ordered chalcopyrite structure, space group  $\bar{1}42d$  (N 122) evolves with composition to a semi-ordered chalcopyrite-like structure,



**Figure 5:** DTA thermograms for  $(\text{CuInSe}_2)_{1-x}(\text{TaSe})_x$  alloys,  $0 \leq x \leq 0.5$ . Left: Heating cycle; Right: Cooling cycle.



**Figure 6:** T-x phase diagram for  $(\text{CuInSe}_2)_{1-x}(\text{TaSe})_x$  alloys,  $0 \leq x \leq 0.5$ . Black circles: heating cycle; White circles: cooling cycle;  $\alpha$ : Ordered chalcopyrite-like phase;  $\alpha'$ : Semi-ordered chalcopyrite phase;  $\beta$ : Disordered zinc-blende phase; L: Liquid phase.

space group  $P\bar{4}2c$  (N 112), until a new cationic rearrangement occurs towards stannite-type structure, space group  $I\bar{4}2m$  (N 121), at  $x = 2/3$  [48,49].

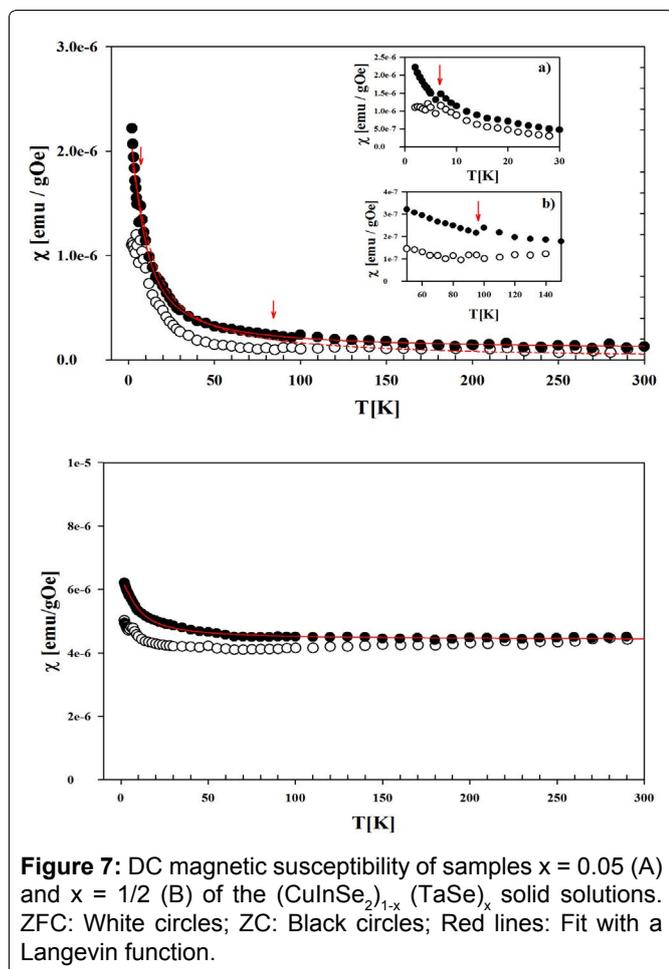
In order to build a preliminary T-x phase diagram in

**Table 4:** Results of the fits for sample  $x = 0.05$  (Figure 6).

Equations	$\mu_m$ [ $\mu_B$ ]	$\chi_0$ [emu/gOe]	$\chi_p$ [emu/gOe]	$R_{sqr}$
$\chi = \chi_0 L(x)$	$3.6 \times 10^3$	$2.1 \times 10^{-6}$	---	0.979
$\chi = \chi_0 L(x) + \chi_p$	$2.4 \times 10^3$	$2.2 \times 10^{-6}$	$8.9 \times 10^{-9}$	0.993

the basis of the information obtained by SEM, XRD and DTA measurements, in Figure 5, DTA thermograms are show. In heating and cooling cycles, a chalcopyrite-like behavior can be clearly observed. Two transitions, one due to cationic sublattice disorder and another due to melting (or solidification), in the composition interval  $0 \leq x \leq 1$ , although sample  $x = 0.05$  show a little peak at 932 K in the heating (907 in the cooling), suggesting that this is the existence field of the ordered chalcopyrite-like ( $\alpha$ ) phase. For  $x > 0.1$ , the thermogram show the evolution to another phase with several thermal transitions. In Figure 6, we show the proposed T-x phase diagram.

The chalcopyrite-like ( $\alpha$ ) phase field is limited to composition range  $0 \leq x \leq 0.1$  at  $T = 800$  K; with increasing temperature the phase field decreases until the temperature of  $T = 1077$  K. The semi-ordered ( $\alpha'$ ) phase exist for compositions  $x > 0.1$ , at  $T = 800$  K until  $T = 932$  K whereas between  $T = 932$  K and  $T = 1077$  K, the  $\alpha'$  phase coexist with the  $\beta$  phase. At  $x = 0$ , the  $\beta$  phase exist between  $T = 1077$  K and the melting point at 1255 K. The melting point decreases between  $x = 0$  to approximately  $x = 0.15$ , then increases until  $x = 0.3$  and decrease again until  $x = 0.5$ . There is a wide liquid + solid region between  $x = 0.15$  and  $x = 0.5$  but region phases cannot be identified.



**Figure 7:** DC magnetic susceptibility of samples  $x = 0.05$  (A) and  $x = 1/2$  (B) of the  $(\text{CuInSe}_2)_{1-x}(\text{TaSe})_x$  solid solutions. ZFC: White circles; ZC: Black circles; Red lines: Fit with a Langevin function.

We also observe overcooling in the samples: the thermal transitions in the cooling cycle are observed at lower temperature than in the heating cycle.

DC magnetic susceptibility measurements were performed in samples  $x = 0.05$  and  $1/2$  using SQUID technique and the ZFC-FC protocol (Figure 7):

Sample  $x = 0.05$  shows a paramagnetic-like behavior with hysteresis between ZFC and FC curves typical of systems with magnetic clusters. The position of two little peaks in the FC curve are signaled by red arrows and amplified in the insets. The origin of these two little peaks cannot be explained unambiguously at this stage of this study; however, a possibility is that it can be related to the blocking-unblocking process of magnetic clusters. Effectively, the peak at lower temperature coincides with a relative maximum of the magnetic susceptibility in the ZFC curve which is typical of systems with a size distribution of the magnetic clusters; this relative maximum in the ZFC curve sometimes also produce a peak in the FC curve as it observed in our measurement. With respect to the little peak at higher temperature, it is observed only in the FC curve and it is very weak, in consequence, their origin remains speculative. Two fits (dashed and contin-

uous red lines) were performed on the FC curve; the first one (dashed line) with a simple Langevin function  $\chi(T) = \chi_0 L(x)$ , where  $x = \mu H/k_B T$ , being  $\mu$  the mean magnetic moment of the magnetic clusters,  $H$  the applied magnetic field (100 Oe in our experiment),  $k_B$  the Boltzmann constant and  $T$  the absolute temperature. It can be observed that this fit is not good for  $T > 100$  K. For the second fit (continuous red line), a constant term ( $\chi_p$ ) has been added to the Langevin function, suggesting the presence of a little magnetic component which can be due to the presence of the  $\text{TaIn}_{0.67}\text{Se}_2$  secondary phase. This time the fit is better (Table 4). The obtained values for the mean magnetic moment show that clusters are composed for  $\sim 10^3$  Ta atoms.

With respect to the sample with  $x = 0.5$ , the magnetic behavior seems to be that of a weak ferromagnetic ( $4.4 \times 10^{-6}$  emu/gOe) with a  $T_c > 300$  K accompanied with a paramagnetic component appreciable only at low temperature. The fit gives a mean magnetic moment of  $2.7 \times 10^3$  Ta-atoms by cluster.

## Conclusions

$(\text{CuInSe}_2)_{1-x}(\text{TaSe})_x$  solid solutions in the composition range  $0 < x \leq 0.5$  have been investigated. XRD shows a tetragonal chalcopyrite-like phase with lattice parameters very close to  $\text{CuInSe}_2$  and the presence of a secondary phase identified as  $\text{TaIn}_{0.67}\text{Se}_2$  that does not belongs to the  $\text{CuInSe}_2$ -TaSe tie line. The analysis of SEM-EDX shows stoichiometric deviations in Cu and Se elements in the order of the experimental error ( $\sim 10\%$ ) for this technique whereas for In and Ta deviations are, on average, of 23.5 and 34%; also, always in average, the contents of Ta in the samples never excess 10% suggesting than this is the solubility limit. This result coincides with DTA and the obtained preliminary T-x phase diagram. As the amount of Ta increases in the samples, the magnetic behavior evolves from paramagnetic to a weak ferromagnetic.

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